

MARKED-UP VERSION OF SUBSTITUTE SPECIFICATION

INORGANIC INTERCALATING NANO-CATALYST FOR THE FIXATION OF CARBON DIOXIDE INTO ALIPHATIC POLYCARBONATE AND A PROCESS FOR PREPARING THE SAME

TECHNICAL FIELD OF THE INVENTION

The present invention relates to an inorganic intercalating nano-catalyst with high activity useful in the synthesis of poly(—alkylene carbonate)s by the copolymerization of carbon dioxide and epoxide.

The present invention also relates to a method for preparing the inorganic intercalating nano-catalyst.

BACKGROUND OF THE INVENTION

Climate warming or the greenhouse effect is mainly attributed to the mass release of carbon dioxide (CO₂) to the atmosphere. It is estimated that carbon dioxide contributes about 66% of climate warming. The CO₂ level in the atmosphere is now reported to be about 345 ppmv (parts per million by volume), and annually increases at a rate of about 1 ppmv due to human activities, ~~especially for the ease of using~~ due to the consumption of mineral fuel. As an effective approach to reduce the release of

greenhouse gas, the utilization of CO₂ has attracted increasing attention recently. Among these, the use of CO₂ as a polymerization monomer is of practical importance. Aliphatic polycarbonates or the block copolymers of polycarbonate and polyether can be prepared via the direct copolymerization of CO₂ with epoxide such as ethylene oxide (EO), propylene oxide (PO), isobutylene oxide (BO), and cyclohexene oxide (CHO). The copolymerization of carbon dioxide with epoxide to produce poly(alkylene carbonate)s was first reported by Inoue and co-workers (Polymer Letters 7, 287(1969); Makromol Chem 130,210(1969)); and described in U.S. Pat. No. 3,585,168. Other processes were described in U.S. Pat. Nos. 3,900,424; 3,953,383; 4,268,684 and 3,248,415 respectively. However, the progress for the commercialization of these poly(alkylene carbonate)s has been very slow, although there are numerous economic advantages associated with the use of an abundant, low cost material like carbon dioxide. The main reason lies with the practical difficulty in preparing large scale organometallic catalysts for commercial usage.

The catalysts reported by Inoue were prepared by reacting diethylzinc with compounds containing active protons, e.g., water, dicarboxylic acids, or dihydric phenols. Typical catalytic activities ranged from 3.8 to 18.9 g polymer / g zinc, and most of the yields fall at the low end of this range. Long polymerization periods of 24 to 48 h were required in order to achieve satisfactory yields and higher molecular weights of the products. It should note that the Inoue catalysts also generated noticeable amounts of by-products or cyclic carbonate and polyether homopolymer that must be removed from the desired polycarbonate products.

Zinc carboxylates have also been described as effective catalysts for CO₂ polymerization. Because zinc carboxylates are stable and safe compounds having no handling problem when compared with diethylzinc, therefore they are promising candidates as practical and commercial catalyst systems. Soga and co-workers reported that the reaction products of zinc hydroxide and aliphatic dicarboxylic acids exhibited high activity for the copolymerization of carbon dioxide and propylene oxide (Polymer J. 13(4), 407(1981)). A variety of acids were investigated, but only adipic and glutaric acid produced the catalysts with higher activity than the known diethylzinc catalysts. Catalysts prepared from aromatic dicarboxylic acids were essentially inert under the polymerization conditions described by Soga.

Soga also reported another approach to improve the catalytic activity by supporting the catalyst on an inert Oxide Carrier (Nippon Kagakkaishi 2, 295(1982)). A supporting material can increase the surface area of active catalyst, thereby enhancing the efficiency production of the aliphatic polycarbonate. However, the supported catalysts of Soga were ineffective compared to the well-known diethylzinc based catalysts.

The metal salts of acetic acid are the third type of catalyst materials known to promote the copolymerization of CO₂ with epoxides (Soga. et al., Makromol. Chem. 178, 893(1977)). Only zinc and cobalt can produce alternating copolymers from CO₂ and epoxides, and the activity of these catalysts remained lower than that derived from diethylzinc catalysts.

In U.S. Pat. No. 4,783,445, Sun reported that soluble zinc catalysts can be prepared by reacting zinc oxide or zinc salts with a dicarboxylic acid anhydride or

monoester in a suitable solvent such as the lower alcohols, ketones, esters and ethers. However, low catalytic activity was produced.

In previous work of U.S. Pat. ~~0,134,740~~ No. 6,844,287, we provided a supported catalyst with high activity useful on the synthesis of poly(-alkylene carbonate)s derived by the copolymerization of CO₂ and epoxides. The highest catalytic activity in that patent is 358.8 g polymer / g zinc. Therefore this kind of catalyst ~~is~~ has the potential for commercial ~~practical~~ use.

Among the catalysts reported in the literature ~~up to that time known to Applicants~~, only zinc carboxylates and supported zinc carboxylates based on adipic or glutaric acid seem to have the potential for practical use on a commercial scale.

BRIEF SUMMARY OF THE INVENTION

One objective of the present invention is to provide a process for preparing an inorganic intercalating or intercalated nano-catalyst comprising intercalating zinc dicarboxylate into the inorganic matrix with a layered structure. The weight ratio of zinc dicarboxylate to inorganic matrix Varies from 1/1 to 1/20.

Another objective of the present invention is to provide an inorganic intercalating or intercalated nano-catalyst obtained by the aforementioned process useful in the copolymerization of carbon dioxide and epoxides to form poly(alkylene carbonate)s.

This invention provides a process for the preparation of an inorganic intercalating nano-catalyst with zinc dicarboxylate as an intercalating agent used for

preparing copolymers from epoxides and carbon dioxide. The zinc dicarboxylates were synthesized from zinc oxide and dicarboxylic acids selected from the group consisting of either succinic acid, glutaric acid, adipic acid, pimelic acid and suberic acid. The inorganic matrices are selected from the group of layered silicates consisting of montmorillonite, mica, vermiculite, kaoline etc. The matrices were activated at 600-1000 DEG C (°C) in a muffle furnace for 2-10 h prior to intercalation. Zinc dicarboxylates ~~was~~ were dissolved in strong polar solvents ~~underhaving~~ under pH values from 1.0 to 4.0. Calcinated acidic matrix was introduced into the reaction system to perform intercalation 30~120 minutes at the temperature from room temperature to 80°C . The solvent was removed and ~~the crystals~~ of the intercalating nano-catalyst ~~was~~were improved by refluxing in ~~less-polar solvent~~ of lower polarity.

In the intercalating process, the strong polar solvents were selected from the group consisting of methanol, glycol, ethylene glycol monobutyl ether, ethylene glycol monomethyl ether, N, N' - dimethyl formamide, sulfolane, imidazole, quinoline, water and N-cyclohexyl pyrrolidine ~~underhaving~~ under pH values from 1.0 to 4.0. Under the help of ~~the strong polar solvents~~ zinc dicarboxylate was introduced into the inter spacing of the interlayers of the inorganic matrix-, ~~T~~thereby increasing the surface area of ~~the~~ the catalysts and enhancing the effective production of the aliphatic polycarbonate.

In the crystal-improving process, the ~~less-polar-solvents~~ of lower polarity were selected from the group consisting of benzene, toluene and xylene, etc. After this process, the ~~crystals~~ of the intercalating agent ~~was~~were improved and redistributed

among the interlayers of the inorganic matrix, thereby increasing the activity of the catalyst and enhancing the effective production of the aliphatic polycarbonate.

The inorganic intercalating nano-catalyst can be used for the copolymerization of CO₂ and PO. The monomers copolymerized in-situ and produced nano-composites. Experimental results showed that extremely highly catalytic activity can be achieved by using zinc glutarate, zinc adipate or zinc pimelate as the intercalating agent.

A more detailed description of the invention and its methods of practice are described in the following examples. It should be understood that the present invention is not intended to limit these examples in any way.

DETAILED DESCRIPTION OF THE INVENTION

Pretreatment of materials

Epoxide, e.g., propylene oxide (PO) with a purity of 99.5% was purified by distillation over calcium hydride under dry nitrogen gas flow for 2 h. The as-treated PO was then stored over 4A molecular sieves prior to use. Carbon dioxide with a purity of higher than 99.8% was used as received. Dicarboxylic acids succinic acid, glutaric acid, adipic acid, pimelic acid and suberic acid were of 98.0% purity, and solvents such as toluene, methanol, acetone, methylene dichloride, were of analytical reagent grade and used without further purification. Zinc oxide of 99.99% purity was also used without further treatment.

Preparation of zinc dicarboxylate

Zinc dicarboxylate was synthesized from zinc oxide and dicarboxylic acid under magnetic stirring as described in the literature. Fine powders of zinc oxide were used as received without further grinding. Accordingly, to a 150 mL ~~containing 90mL~~ toluene-three-neck round bottom flask containing 90mL toluene and equipped with a magnetic stirrer, condenser, and a Dean-Stark trap ~~were~~was added a slight molar excess of zinc oxide. To this mixture ~~were~~was then introduced dicarboxylic acid, and the mixture was slowly heated up to 55-110°C for 4 to 20 h under vigorous stirring. Upon cooling down, the resulting mixture was filtered. The resulting solids were continuously washed with acetone for several times followed by drying overnight in a vacuum oven at 100°C. The obtained zinc dicarboxylates were fine powders in white color with a high acid conversion of $\geq 99\%$.

Preparation of inorganic intercalating nano-catalysts

The preparation process of inorganic intercalating nano-catalysts involved three steps. Firstly, the layered matrices were delaminated with diluted acid, then calcined at 600-1000 °C in a muffle furnace for 2~10 h prior to intercalation. Subsequently, zinc dicarboxylates were dissolved in strong polar solvents ~~under~~having pH values from 1.0 to 4.0, then calcinated acidic matrices were introduced into the reaction system to perform intercalation for 30~120 minutes at the temperature from room temperature to 80°C. Herein the strong polar solvents are selected from methanol, glycol, ethylene glycol monobutyl ether, ethylene glycol monomethyl ether, N, N'-dimethyl formamide, sulfolane, imidazole, quinoline, water and N-cyclohexyl

pyrrolidine. Finally, the solvents were removed and the crystals of the intercalating nano-catalysts were improved by refluxing in less polar solvents. Herein the less polar lower polarity solvents are selected from benzene, toluene, xylene, etc. The intercalating process was performed in a 150 mL single-neck round bottom flask equipped with a magnetic stirrer, and a condenser. The crystal-improving process was carried out with a set of refluxing equipment.

Copolymerization

The copolymerization of CO₂ and epoxide, e.g. propylene oxide was carried out in a 500mL autoclave equipped with a mechanical stirrer. Inorganic intercalating nano-catalyst was further dried at 100°C for 24-h prior to being used for the polymerization process. Dry inorganic intercalating nano-catalyst was then introduced into the autoclave as quickly as possible. The autoclave was then capped with its head, and the entire assembly was connected to the reaction system equipped with a vacuum line. The autoclave with catalyst inside was further dried for 24h under vacuum at 100°C. This implied that the catalyst was further in-situ dried during the same process for another 24-h. Subsequently, the autoclave was purged with carbon dioxide and evacuated alternatively for three times, followed by adding purified PO with a syringe. The autoclave was then pressurized to 5.0 MPa via a CO₂ cylinder. The copolymerization was performed at 60°C under stirring for 40-h. The autoclave was cooled to room temperature and the pressure was released. The resulting viscous mixture was removed and poured into vigorously stirred methanol to precipitate poly (propylene carbonate) (PPC). The as-made PPC was filtered and dried for two days at room

temperature under vacuum. _Meanwhile, the resulting filtrate was distilled to remove methanol and methylene chloride to yield a methanol soluble product.

The catalytic activity was calculated by dividing the mass of the yielded PPC with the mass of zinc containing in the said intercalating nano-catalyst.

Example 1

Pretreatment of vermiculite

The vermiculite (VMT) was first pretreated with hydrochloric acid as described in the literature. _To a 1L polypropylene beaker containing 800 mL of 2M HCl solution was added 25.0 g of 250 mesh crude VMT at room temperature. _The resulting slurry was magnetically stirred for 12-h. _The VMT was separated by filtration and then washed thoroughly with distilled water several times until the filtrate had a pH value of 7.0. _After the washing—, the obtained solid material was dried at 300°C over ~~night~~overnight then calcinated at 1000 °C for 4-h.

Intercalation reaction

1.0 g zinc succinate was dissolved in a 150mL flask containing 100mL water with a pH value of 2.0, then calcinated acidic VMT 3.0 g was introduced into the solution to perform intercalation at room temperature for 30 minutes. Next,~~The removed~~ the solvent of this system was removed and ~~obtained~~ crude inorganic intercalating nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced to~~The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL benzene and ~~kept on~~-refluxing

was maintained at 80° C for 24-h. Then the final catalyst was separated by filtration and dried at 100°C overnight.

Evaluation of catalytic activity

~~Carefully, weighed 2.0 g so made of catalyst made as above was carefully weighed and performed copolymerization was performed as described above.~~ The catalytic activity of this catalyst is 10.4 g polymer / g zinc.

Examples 2-7

Pretreatment of vermiculite

VMT was first pretreated with hydrochloric acid as described in the literature-. To a 2L polypropylene beaker containing 1600 mL of 2M HCl solution was added 50.0 g of 250 mesh crude VMT at room temperature. The resulting slurry was magnetically stirred for 12-h. The VMT was separated by filtration and then washed thoroughly with distilled water several times until the filtrate had a pH value of 7.0. After the washing-, the obtained solid material was dried at 300°C ~~over night~~ overnight then calcinated at 900°C for 4-h.

Example 2

Intercalation reaction

1.0 g zinc glutarate was dissolved in a 150mL flask containing 100mL water with a pH value of 3.0, then calcinated -acidic VMT 1.0 g was introduced into the solution to perform intercalation at room temperature for 60 minutes. ~~Then removed~~

Next, the solvent of this system was removed and obtained crude inorganic intercalating nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with, refluxing equipment containing 300 mL toluene and kept on refluxing was maintained at 110°C for 24-h. Then the final catalyst was separated by filtration and dried at 100°C overnight.

Evaluation of catalytic activity

~~Carefully weighed 2.0 g so-made of catalyst~~ made as above was carefully weighed and performed copolymerization was performed as described above. The catalytic activity of this catalyst is 400.5 g polymer / g zinc.

Example 3

Intercalation reaction

1.0 g zinc glutarate was dissolved in a 150mL flask containing 100mL water with a pH value of 3.0, then calcinated acidic VMT 2.0 g was introduced into the solution to perform intercalation at room temperature for 60 minutes. ~~Then removed~~ The solvent of this system was removed and obtained crude inorganic intercalating nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL toluene and kept on refluxing was

maintained at 110°C for 24-h. Then the final catalyst was separated by filtration and
dried at 100°C overnight.

Evaluation of catalytic activity

~~Carefully weighed 2.0 g so made of catalyst made as above was carefully~~
weighed and ~~performed copolymerization~~ was performed as described above. The
catalytic activity of this catalyst is 405.8 g polymer / g zinc.

Example 4

Intercalation reaction

1.0 g zinc glutarate was dissolved in a 150mL flask containing 100mL
water with a pH value of 3.0, then calcinated acidic VMT 3.0 g was introduced into the
solution to perform intercalation at room, temperature for 60 minutes. ~~Then removed~~
Next, the solvent of this system was removed and obtained crude inorganic intercalating
nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500
mL flask with refluxing equipment containing 300 mL toluene and ~~kept on refluxing~~ was
maintained at 110°C for 24-h. Then the final catalyst was separated by filtration and
dried at 100°C overnight.

Evaluation of catalytic activity

~~Carefully weighed 2.0 g so-made~~ of catalyst made as above was carefully weighed and ~~performed copolymerizaion~~ was performed as described above. The catalytic activity of this catalyst is 410.8 g polymer / g zinc.

Example 5

Intercalation reaction

1.0 g zinc glutarate was dissolved in a 150mL flask containing 100mL water with a pH value of 3.0, then calcinated acidic VMT 5.0 g was introduced into the solution to perform intercalation at room temperature for 60 minutes. ~~Then removed~~ Next, the solvent of this system was removed and ~~obtained crude inorganic intercalating~~ nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL toluene and ~~kept on refluxing~~ was maintained at 110° C for 24-h. Then the final catalyst was separated by filtration and dried at 100° C overnight.

Evaluation of catalytic activity

~~Carefully weighed 2.0 g so-made~~ of catalyst made as above was carefully weighed and ~~performed copolymerization~~ was performed, as described above. The catalytic activity of this catalyst is 420.8 g polymer / g zinc.

Example 6

Intercalation reaction

1.0 g zinc glutarate was dissolved in a 150mL flask containing 100mL water with a pH value of 3.0, then calcinated acidic VMT 10.0g was introduced into the solution to perform intercalation at room temperature for 120 minutes. ~~Then removed~~ Next, the solvent of this system was removed and obtained crude inorganic intercalating nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL toluene and kept on refluxing was maintained at 110° C for 24-h. Then the final catalyst was separated by filtration and dried at 100° C overnight.

Evaluation of catalytic activity

~~Carefully weighed 4.0 g so made of catalyst made as above was carefully weighed and performed copolymerization as described above was performed.~~ The catalytic activity of this catalyst is 415.4 g polymer / g zinc.

Example 7

Intercalation reaction

1.0 g zinc glutarate was dissolved in a 150mL flask containing 100mL water with a pH value of 3.0, then calcinated acidic VMT 20.0g was introduced into the solution to perform intercalation at room temperature for 120 minutes. ~~Then removed~~

Next, the solvent of this system was removed and obtained crude inorganic intercalating nano-catalyst with VMT as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL toluene and ~~kept on refluxing~~ was maintained at 110°C for 24-h. Then the final catalyst was separated by filtration and dried at 100°C overnight.

Evaluation of catalytic activity

~~Weighed 5.0 g so made of catalyst made as above was weighed out with an~~ analytic balance and performed copolymerization was performed as described above. The catalytic activity of this catalyst is 403.8 g polymer / g zinc.

Example 8

Pretreatment of Montorillonite

~~The M~~ montorillonite was first pretreated with hydrochloric acid as described in the literature. To a 1L polypropylene beaker containing 800 mL of 2M HCl solution was added 25.0 g of 250 mesh crude montorillonite at room temperature. The resulting slurry was magnetically stirred for 12-h. The montorillonite was separated by filtration and then washed thoroughly with distilled water several times until the filtrate had a pH value of 7.0. After the washing, the obtained solid material was dried at 300°C ~~over night~~ overnight then calcinated at 800°C for 2-h.

Intercalation reaction

1.0 g zinc adipate was dissolved in a 500mL flask containing 400mL water with a pH value of 4.0, then calcinated acidic montorillonite 3.0 g was introduced into the solution to perform intercalation at 60°C for 60 minutes. ~~Then removed~~ Next, the solvent of this system was removed and ~~obtained~~ crude inorganic intercalating nano-catalyst with montorillonite as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL xylene and ~~kept on refluxing~~ was maintained at 138°C for 24-h. Then the final catalyst was separated by filtration and dried at 100°C overnight.

Evaluation of catalytic activity

~~Carefully weighed~~ 2.0 g ~~so made of~~ catalyst made as above was carefully weighed and ~~performed copolymerization~~ was performed as described above. The catalytic activity of this catalyst is 339.4 g polymer / g zinc.

Example 9

Pretreatment of mica

The mica was first pretreated with hydrochloric acid as described in the literature. To a 1L polypropylene beaker containing 800 mL of 2M HCl solution was added 25.0 g of 250 mesh crude mica at room temperature. The resulting slurry was magnetically stirred for 12-h. The mica was separated by filtration and then washed

thoroughly with distilled water several times until the filtrate had a pH value of 7.0. After the washing, the obtained solid material was dried at 300 °C ~~over night~~ overnight then calcinated at 750 °C for 2-h.

Intercalation reaction

1.0 g zinc pimelate was dissolved in a 500mL flask containing 400mL water with a pH value of 4.0, then calcinated acidic mica 3.0 g was introduced into the solution to perform intercalation at 80°-C for 120 minutes. ~~Then removed~~ Next, the solvent of this system was removed and obtained crude inorganic intercalating nano-catalyst with mica as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL toluene and kept on refluxing was maintained at 110° C for 24h. Then the final catalyst was separated by filtration and dried at 100° C overnight.

Evaluation of catalytic activity

~~Carefully weighed 2.0g so made of catalyst~~ made as above was weighed and performed copolymerization was performed as described above. The catalytic activity of this catalyst is 240.9 g polymer / g zinc.

Example 10

Pretreatment of kaoline

~~The K~~ kaoline was first pretreated with hydrochloric acid as described in the literature. To a 1L polypropylene beaker containing 800 mL of 2M HCl solution was

added 25.0 g of 250 mesh crude kaoline at room temperature. The resulting slurry was magnetically stirred for 12-h. The kaoline was separated by filtration and then washed thoroughly with distilled water several times until the filtrate had a pH value of 7.0. After the washing, the obtained solid material was dried at 300°C ~~over night~~ overnight then calcinated at 600°C for 4-h.

Intercalation reaction

1.0 g zinc suberate was dissolved in a 500mL flask containing 400mL water with a pH value of 4.0, then calcinated acidic kaoline 3.0 g was introduced into the solution to perform intercalation at 60°C for 60 minutes. ~~Then removed~~ Next, the solvent of this system was removed and ~~obtained~~ crude inorganic intercalating nano-catalyst with kaoline as matrix was obtained.

Crystal improving process

~~Introduced~~ The above obtained crude catalyst was introduced into a 500 mL flask with refluxing equipment containing 300 mL benzene and ~~kept on refluxing~~ was maintained at 80°C for 24-h. Then the final catalyst was separated by filtration and dried at 100°C overnight.

Evaluation of catalytic activity

~~Carefully weighed 2.0 g so made of catalyst~~ was carefully weighed and ~~performed copolymerization~~ was performed as described above. The catalytic activity of this catalyst is 128.8 g polymer / g zinc.

Table 1. Summary of the inorganic intercalating nano-catalysts and their catalytic activity

Example	Intercalating agent	Inorganic matrix	Matrix content (%) [*]	Calcination condition (°C/hr)	Intercalation condition	Crystal improving conditions	Catalytic activity (g polymer/g zinc)**
1	Zinc succinate	Vermiculite,	300	1000°C/4	Rom temperature / 30 min	benzene, 80 °C	10.4
2	Zinc glutarate	Vermiculite	100	900°C/4	Rom temperature / 60 min	toulene, 110 °C	400.5
3	Zinc glutarate	Vermiculite	200	900°C/4	Rom temperature / 60 min	toulene, 110 °C	405.8
4	Zinc glutarate	Vermiculite	300	900°C/4	Rom temperature / 60 min	toulene, 110 °C	410.8
5	Zinc glutarate	Vermiculite	500	900°C/4	Rom temperature / 60 min	Toulene, 110 °C	420.8
6	Zinc glutarate	Vermiculite	1000	900°C/4	Rom temperature / 120min	toulene, 110 °C	415.4
7	Zinc glutarate	Vermiculite	2000	900°C/4	Rom temperature / 120min	toulene, 110 °C	403.8
8	Zinc adipate	Montmorillonite	300	800°C/2	60 °C / 60 min	xylene, 138 °C	339.4
9	Zinc pimelate	mica	300	750°C/2	80 °C / 120 min	toulene, 110 °C	240.9
10	Zinc suberate	kaoline	300	600°C/4	60 °C / 60 min	Benezene, 80 °C	128.8

**copolymerization condition: 60 °C, 40 h, CO₂ pressure: 5.2MPa.

*Inorganic content calculated based on intercalating agent.